

---

# EFFECT OF SYNTHESIS TIME ON FORMATION OF $\text{Fe}_3\text{O}_4$ NANOPARTICLES USING THE SOLVOTHERMAL PROCESS

ELAINE YANG

REU SYMPOSIUM 8/10/18 DR JIAO'S LAB

# OVERVIEW

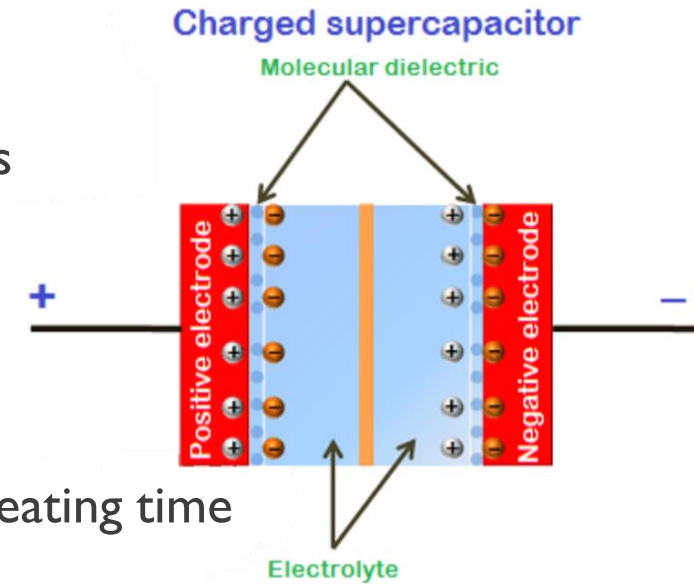
- Introduction
- Method
- Results
- Discussion
- Conclusion

# INTRODUCTION

- Transition oxide nanoparticles (NPs) exhibit high magnetic signals and can be altered by external magnetic fields
  - Unique physical and chemical properties; lighter and stronger; high surface area-volume ratio
  - Magnetite ( $\text{Fe}_3\text{O}_4$ ) NPs are ferrimagnetic
  - Very abundant, environmentally friendly, and can be “grown” rather than built
  - Functional groups can be used for more sensitive biolabeling, biodelivery, and higher contrast MRIs
- Hybridized graphene and iron oxide NPs have improved specific capacitance and less electrical resistance
  - Graphene has a high surface area and electrical conductivity
    - Can serve as substrate and conductive pathway for NPs

# INTRODUCTION CONT.

- Hybrid used as electrode material in supercapacitors/double-layer capacitors
  - More energy efficient than traditional batteries and capacitors
  - Builds up charges on two metal plates to store static electricity
  - Creates thin double-layer of charge
- Properties of NPs altered by precursor material, heating temperature, and heating time
- Size, shape, and distribution largely impact its magnetic properties
- One-step procedure produces NPs attached to the graphene surface
  - Hypothesized that nuclei are absorbed into spherical NPs through intermolecular forces
- **Purpose:** to explore the effect of synthesis time on altering the size and distribution of iron oxide nanoparticles on graphene sheets



# METHOD OVERVIEW

- Iron (III) acetylacetonate, expanded graphite, ethanol
- Simple solvothermal process forms graphene attached to magnetite nanoparticles, which can be used for other transition metals
- **Graphene Exfoliation**
  - Sonicate expandable graphite with NMP and centrifuge graphene out
- **Solvothermal Synthesis**
  - Place iron (III) acetylacetonate and 200 proof ethanol in an autoclave
  - Heat autoclave in a muffle furnace at 180 C for varying amounts of time
  - Put autoclave under forced air cooling and centrifuge product out
- Test various heating times to see effects on particle size and coating density



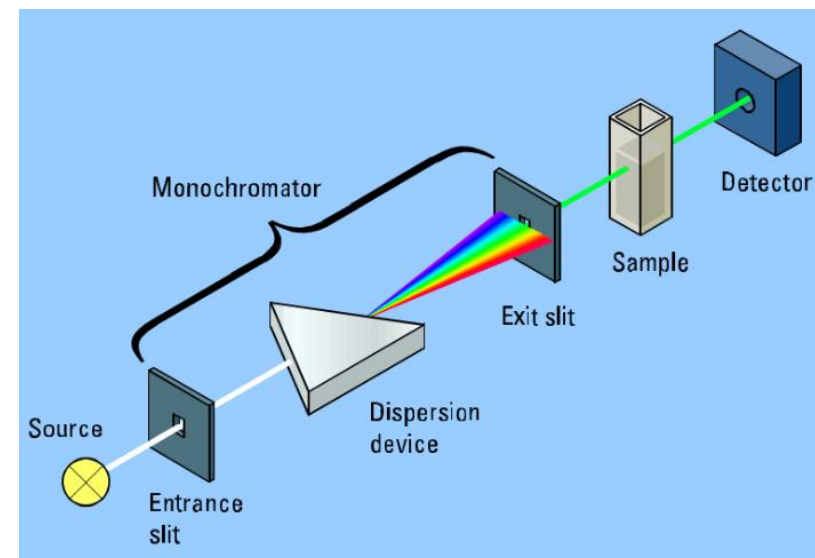
# CHARACTERIZATION

## ■ UV-Vis Spectroscopy

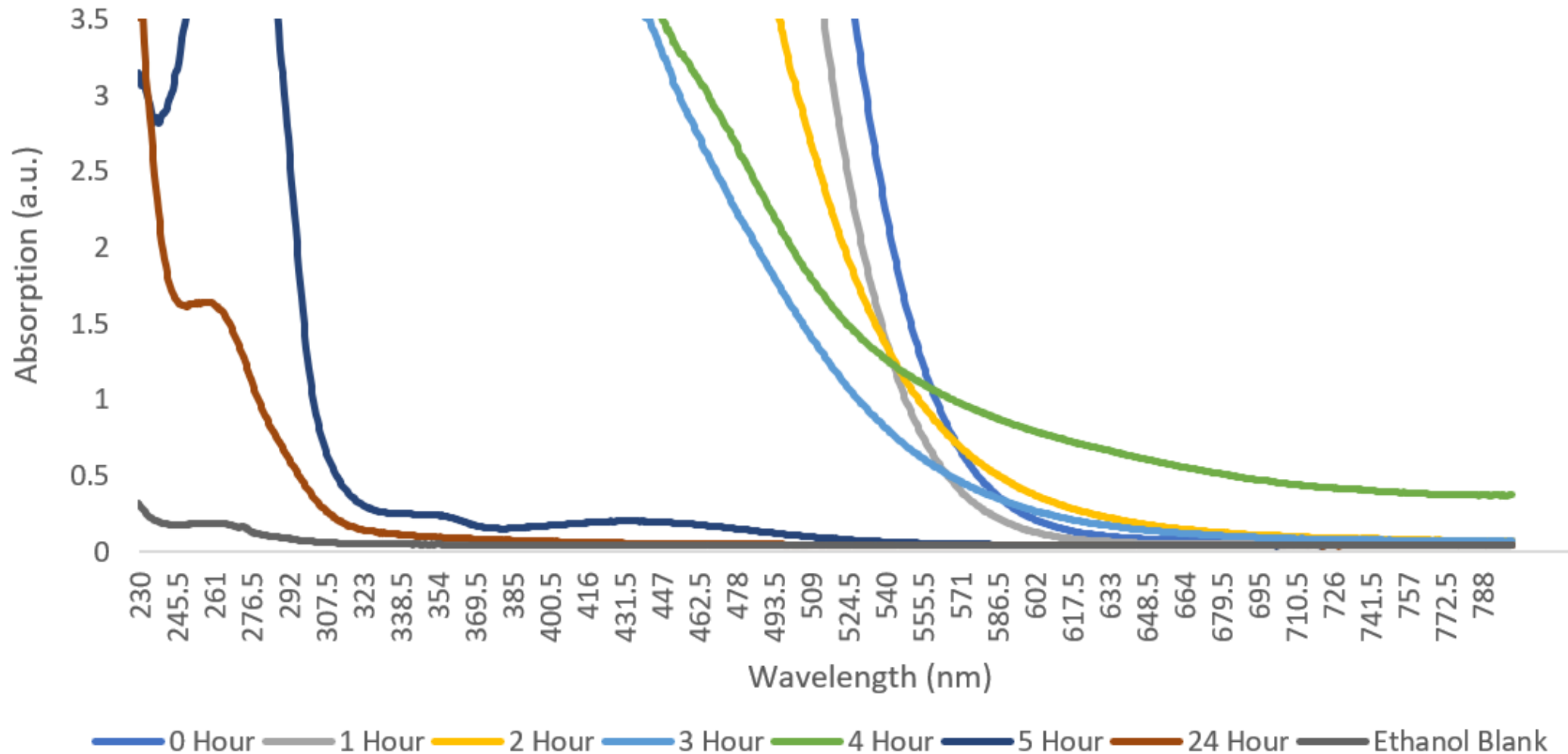
- Remaining solution placed into a cuvette
- Energy is passed through and absorbed to excite electrons to higher energy orbitals
- Detector records degree of absorption at each wavelength
- Beer-Lambert Law states absorbance is proportional to the concentration of a solution
- Measures change in concentration of reactants or products over time

## ■ Transmission Electron Microscope (TEM)

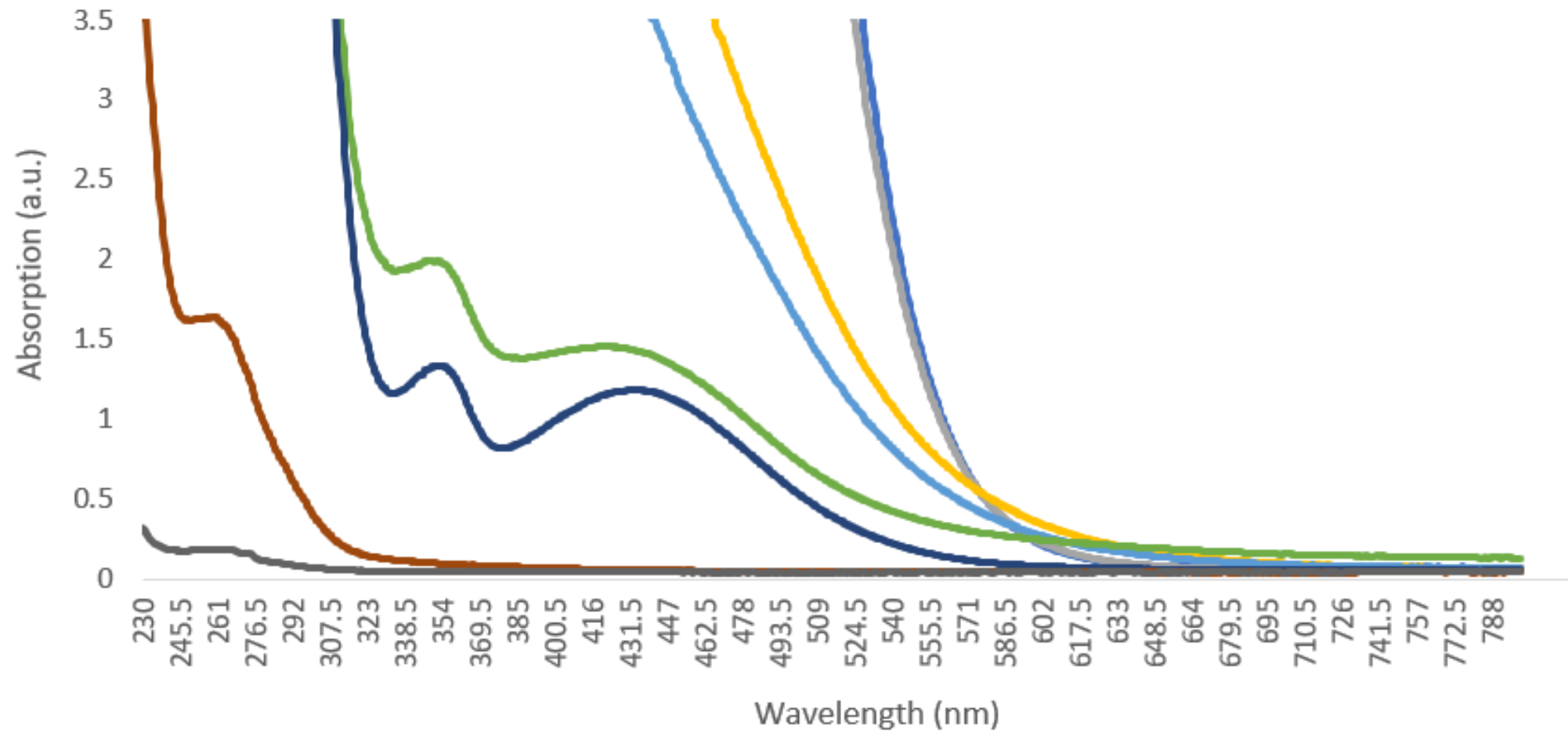
- Used to observe and characterize possible structures that are formed



# UV-VIS OF HEATING TIME VS ABSORPTION (TRIAL I)

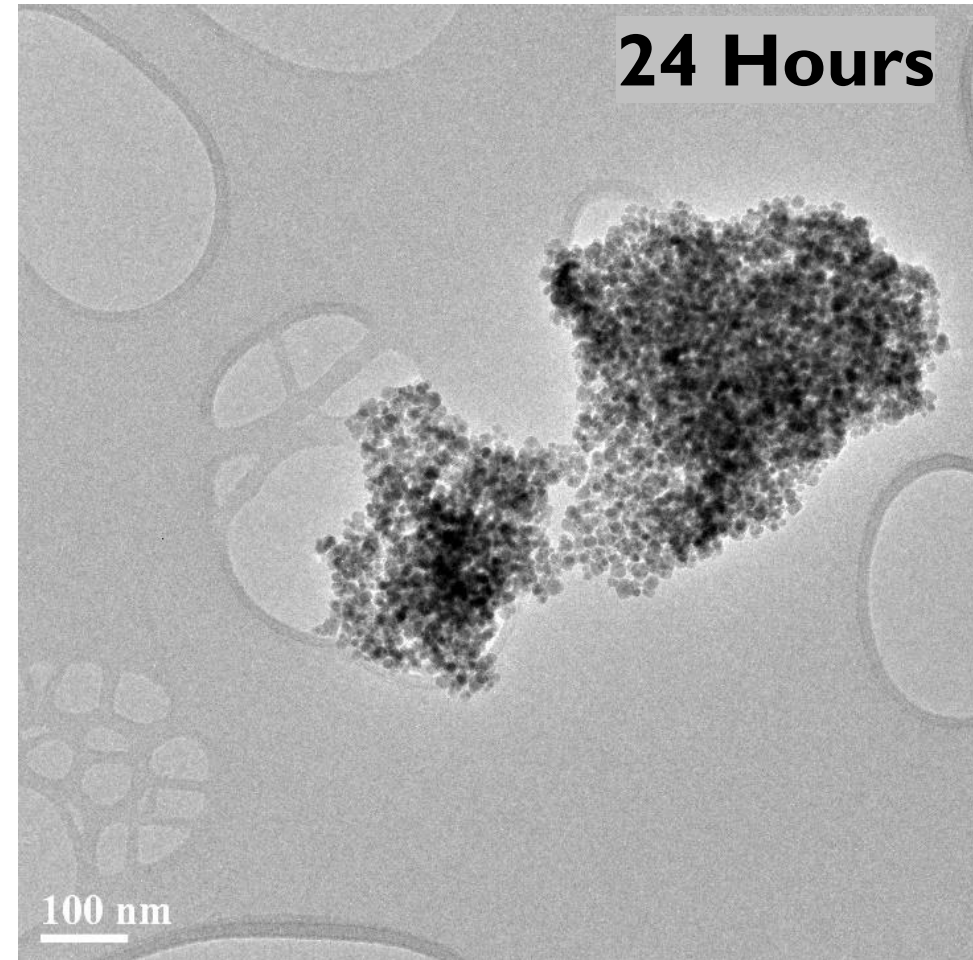
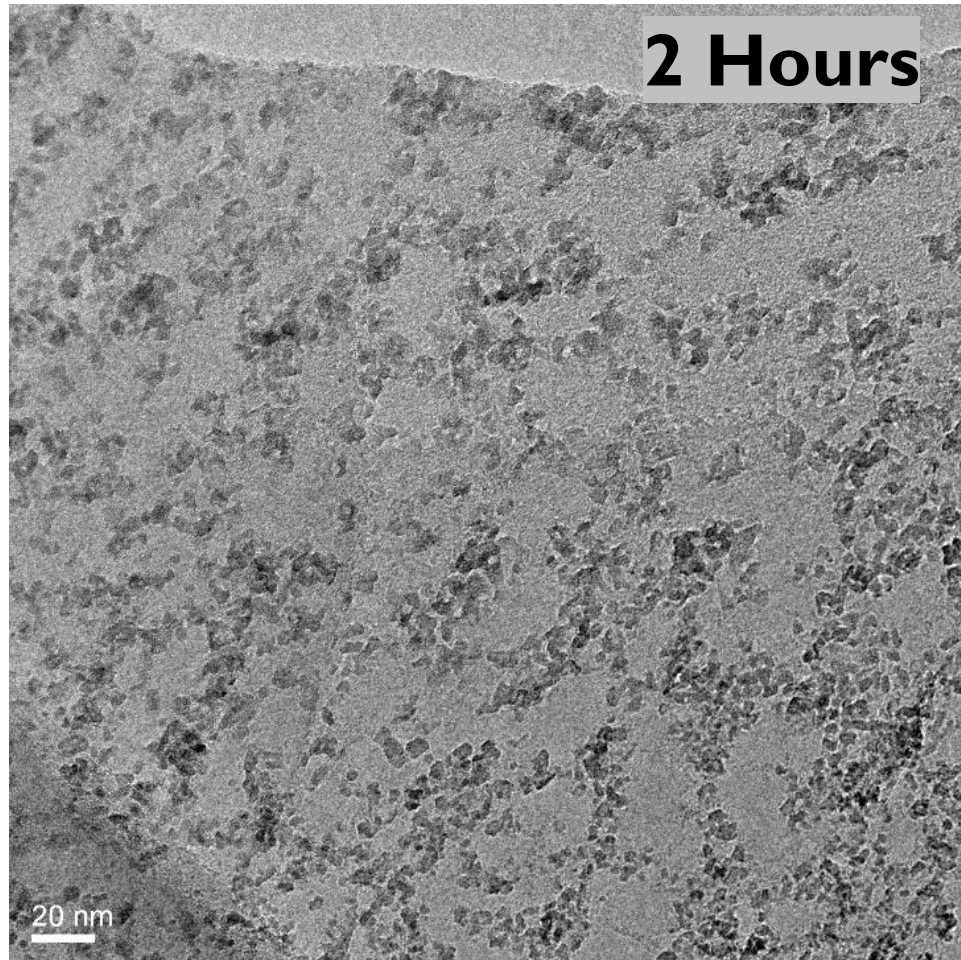


# UV-VIS OF HEATING TIME VS ABSORPTION (TRIAL 2)





# TEM IMAGES OF $\text{Fe}_3\text{O}_4$ NANOPARTICLES



# DISCUSSION

## ■ **UV-Vis Spectroscopy**

- First trial demonstrated a sudden jump between 4 and 5 hours of heating
- Repeated a second trial due to leaks during the synthesis process and unclear results
- Second trial showed a less dramatic jump between 3 and 4 hours of heating
- Most of the synthesis completes within 5 hours of heating

## ■ **Transmission Electron Microscope (TEM)**

- Nanoparticles were being formed on the carbon support even with only 2 hours of heating
- Synthesis is able to proceed (but not complete) within a short amount of time

# CONCLUSION

- Majority of the reaction completes within 5 hours of heating time, which is much shorter than the previously hypothesized 16 hours
  - Allows for more efficiency with reduced synthesis times of iron oxide nanoparticles
- Understanding the growth pattern of the nanoparticles allows us to more accurately tailor the particle size and distribution
- Future work would center around gathering consistent data on heating times from 0 to 24 hours
- Can be used for large-scale production and more environmentally friendly devices due to its wide range of capabilities and possible applications

## ACKNOWLEDGEMENTS

- This research was made possible by the generous support of the National Science Foundation to fund the REU program and Saturday Academy for providing this opportunity. I would like to thank my mentors and reviewers from PSU who guided me through the research process and provided me with a space to conduct my project. I would also like to thank my lab partners for their constant support and assistance.





# REFERENCES

- [1] A. Ali *et al.*, “Synthesis, characterization, applications, and challenges of iron oxide nanoparticles,” *Nanotechnol. Sci. Appl.*, vol. 9, pp. 49–67, 2016.
- [2] W. Qian *et al.*, “Surfactant-free hybridization of transition metal oxide nanoparticles with conductive graphene for high-performance supercapacitor,” *Green Chem.*, vol. 14, no. 2, pp. 371–377, 2012.
- [3] J. Xie, S. Peng, N. Brower, N. Pourmand, S. X. Wang, and S. Sun, “One-pot synthesis of monodisperse iron oxide nanoparticles for potential biomedical applications,” *Pure Appl. Chem.*, vol. 78, no. 5, pp. 1003–1014, 2006.
- [4] W. Wu, Q. He, and C. Jiang, “Magnetic iron oxide nanoparticles: Synthesis and surface functionalization strategies,” *Nanoscale Res. Lett.*, vol. 3, no. 11, pp. 397–415, 2008.
- [5] L. Zhuang, W. Zhang, Y. Zhao, H. Shen, H. Lin, and J. Liang, “Preparation and characterization of Fe<sub>3</sub>O<sub>4</sub> particles with novel nanosheets morphology and magneto-chromatic property by a modified solvothermal method,” *Sci. Rep.*, vol. 5, pp. 1–6, 2015.
- [6] Z. Kozakova *et al.*, “The formation mechanism of iron oxide nanoparticles within the microwave-assisted solvothermal synthesis and its correlation with the structural and magnetic properties,” *Dalt. Trans.*, vol. 44, no. 48, pp. 21099–21108, 2015.
- [7] W. Reusch, “Visible and Ultraviolet Spectroscopy,” *IOCD*, 2013. [Online]. Available: <https://www2.chemistry.msu.edu/faculty/reusch/virttxtjml/spectrpy/uv-vis/spectrum.htm>. [Accessed: 03-Aug-2018].
- [8] “Introduction to Ultraviolet-Visible Spectroscopy (UV),” *Royal Society of Chemistry*, 2009. [Online]. Available: [http://www.rsc.org/learn-chemistry/content/filerepository/CMP/00/001/304/UV-Vis\\_Student\\_resource\\_pack\\_ENGLISH.pdf](http://www.rsc.org/learn-chemistry/content/filerepository/CMP/00/001/304/UV-Vis_Student_resource_pack_ENGLISH.pdf). [Accessed: 03-Aug-2018].
- [9] S. T. Picraux, “Nanotechnology,” *Encyclopaedia Britannica*, 2018. [Online]. Available: <https://www.britannica.com/technology/nanotechnology>. [Accessed: 11-Jul-2018].
- [https://www.google.com/url?sa=i&source=images&cd=&cad=rja&uact=8&ved=2ahUKEwxi6rMs9\\_cAhXI6oMKHUUCDz8QjRx6BAgBEAU&url=http%3A%2F%2Fwww.physics-and-radio-electronics.com%2Felectronic-devices-and-circuits%2Fpassive-components%2Fcapacitors%2Fsupercapacitor.html&psig=AOvVaw3HjI17B8-e92GrD3n7V\\_I&t&ust=1533884351142905](https://www.google.com/url?sa=i&source=images&cd=&cad=rja&uact=8&ved=2ahUKEwxi6rMs9_cAhXI6oMKHUUCDz8QjRx6BAgBEAU&url=http%3A%2F%2Fwww.physics-and-radio-electronics.com%2Felectronic-devices-and-circuits%2Fpassive-components%2Fcapacitors%2Fsupercapacitor.html&psig=AOvVaw3HjI17B8-e92GrD3n7V_I&t&ust=1533884351142905)
- [http://faculty.sdmiramar.edu/fgarces/labmatters/instruments/uv\\_vis/Cary50\\_Pic/conventional-spectrophotometer.png](http://faculty.sdmiramar.edu/fgarces/labmatters/instruments/uv_vis/Cary50_Pic/conventional-spectrophotometer.png)